

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

3-(2-Acetylanilino)propanoic acid

Christopher R. Sparrow,^a Edwin H. Walker Jr^a and Frank R. Fronczek^{b*}^aDepartment of Chemistry, Southern University, Baton Rouge, LA 70813, USA, and^bDepartment of Chemistry, Louisiana State University, Baton Rouge, LA 70803-1804, USA

Correspondence e-mail: ffroncz@lsu.edu

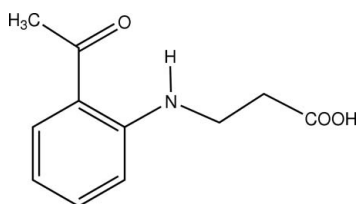
Received 23 October 2008; accepted 29 October 2008

Key indicators: single-crystal X-ray study; $T = 90$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.041; wR factor = 0.113; data-to-parameter ratio = 21.1.

The title molecule, $\text{C}_{11}\text{H}_{13}\text{NO}_3$, has its propanoic acid group in an extended conformation, such that the molecule is nearly planar, with a mean deviation of 0.036 Å [the maxima being 0.106 (1) and 0.110 (1) Å for the two methylene C atoms]. The NH group forms an intramolecular hydrogen bond with the acetyl group; in the crystal COOH group forms a centrosymmetric hydrogen-bonded dimer.

Related literature

For general background, see: Crosby *et al.* (1961, 1962); Foley *et al.* (2003); Walker Jr *et al.* (2004); Yoshihara *et al.* (2001). For related structures, see: Slater *et al.* (2006). For hydrogen-bonding patterns, see: Etter (1990).



Experimental

Crystal data

$\text{C}_{11}\text{H}_{13}\text{NO}_3$ $\gamma = 83.019$ (12)°
 $M_r = 207.22$ $V = 489.32$ (16) Å³
 Triclinic, $P\bar{1}$ $Z = 2$
 $a = 5.1935$ (10) Å Mo $K\alpha$ radiation
 $b = 9.8342$ (16) Å $\mu = 0.10$ mm⁻¹
 $c = 9.920$ (2) Å $T = 90.0$ (5) K
 $\alpha = 77.084$ (12)° $0.30 \times 0.20 \times 0.10$ mm
 $\beta = 85.174$ (11)°

Data collection

Nonius KappaCCD diffractometer 11014 measured reflections
 with an Oxford Cryosystems 3012 independent reflections
 Cryostream cooler 2467 reflections with $I > 2\sigma(I)$
 Absorption correction: none $R_{\text{int}} = 0.021$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.041$
 $wR(F^2) = 0.113$
 $S = 1.04$
 3012 reflections
 143 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.39$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.29$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O2}-\text{H2O}\cdots\text{O3}^i$	0.909 (16)	1.747 (16)	2.6531 (11)	174.7 (14)
$\text{N1}-\text{H1N}\cdots\text{O1}$	0.887 (14)	1.980 (14)	2.6690 (12)	133.4 (11)

Symmetry code: (i) $-x, -y, -z + 2$.

Data collection: COLLECT (Nonius, 2000); cell refinement: SCALEPACK (Otwinowski & Minor, 1997); data reduction: DENZO (Otwinowski & Minor, 1997) and SCALEPACK; program(s) used to solve structure: SIR97 (Altomare *et al.*, 1999); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997); software used to prepare material for publication: SHELXL97.

This research was made possible by a grant supplied by the National Science Foundation's Early CAREER program (Cooperative Agreement DMR-0449886) at Southern University, and by a grant supplied by the US Department of Education Title III Part B HBGI program (grant No. P031B040030) at Southern University. The purchase of the NMR was made possible by the National Science Foundation's Major Research Instrument Program (Cooperative Agreement CHE-0321591) at Southern University. The purchase of the FTIR was made possible by grant No. LEQSF(2005-2007)-ENH-TR-65, and the purchase of the diffractometer was made possible by grant No. LEQSF(1999-2000)-ENH-TR-13, administered by the Louisiana Board of Regents.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: PV2116).

References

- Altomare, A., Burla, M. C., Camalli, M., Cascarano, G. L., Giacovazzo, C., Guagliardi, A., Moliterni, A. G. G., Polidori, G. & Spagna, R. (1999). *J. Appl. Cryst.* **32**, 115-119.
- Crosby, G. A., Alire, R. M. & Whan, R. E. (1961). *J. Chem. Phys.* **34**, 743-748.
- Crosby, G. A., Whan, R. E. & Freeman, J. J. (1962). *J. Chem. Phys.* **66**, 2493-2499.
- Etter, M. C. (1990). *Acc. Chem. Res.* **23**, 120-126.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Foley, T. J., Harrison, B. S., Kneely, A. S., Abboud, K. A., Reynolds, J. R., Schanze, K. S. & Boncella, J. M. (2003). *Inorg. Chem.* **42**, 5023-5032.
- Nonius (2000). COLLECT. Nonius BV, Delft, The Netherlands.
- Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography*, Part A, edited by C. W. Carter Jr & R. M. Sweet, pp. 307-326. New York: Academic Press.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112-122.
- Slater, H. L., Rozynski, H., Crundwell, G. & Glagovich, N. M. (2006). *Acta Cryst.* **E62**, o1957-o1958.
- Walker, E. H. Jr, Apblett, A. W., Walker, R. & Zachary, A. (2004). *Chem. Mater.* **16**, 5336-5343.
- Yoshihara, T., Shimada, H., Shizuka, H. & Tobita, S. (2001). *Phys. Chem. Chem. Phys.* **3**, 4972-4978.

supplementary materials

Acta Cryst. (2008). E64, o2359 [doi:10.1107/S1600536808035277]

3-(2-Acetylanilino)propanoic acid

C. R. Sparrow, E. H. Walker Jr and F. R. Fronczek

Comment

Since the discovery that energy transfer from the triplet state of an organic ligand can efficiently sensitize the emissive states of metal ions (Crosby *et al.*, 1961; 1962) there has been considerable effort devoted to designing ligands that optimize this energy transfer and thus give efficient metal luminescence (Foley *et al.*, 2003). The radiationless energy transfer photoluminescence and/or electro-luminescence properties of aromatic carbonyl compounds are strongly affected by the presence of a substituent on the aromatic ring. In general, the π, π^* state is stabilized by introducing an electron-donating substituent on the aromatic ring, while the location of the n, π^* state is only slightly modified by the electron-donating substituent. Thus, the fluorescence properties of the compounds with close-lying $^1(n, \pi^*)$ and $^1(\pi, \pi^*)$ states are expected to depend markedly on the nature of the solvent, such as its polarity and hydrogen-bonding ability. In 2'-aminoacetophenone, the $^1(n, \pi^*)$ and $^1(\pi, \pi^*)$ states are closely located in the lowest excited singlet state because of the presence of a strong electron-donating substituent. Owing to the proximity of two electronic levels, the photophysical properties of 2'-aminoacetophenone is very sensitive to environment, such as solvent polarity and temperature (Yoshihara *et al.*, 2001). Therefore, the addition of a strong electron-donating substituent such as acrylic acid, which has the ability to form a solventless gel *via* hydrogen bonding, has been synthesized. The structure of the title compound, also known as 3-[(*N*-acetyl-phenyl)-azanediyl]-propionic acid, is herein described. Its synthesis takes advantage of the self-initiating condensation of 2'-aminoacetophenone with the vinyl group of the *ab* unsaturated acrylic acid *via* anti-Markovnikov addition, which is similar to chemistry involved in the synthesis of the novel 3,3',3''-nitriлотripropionic acid precursor gel that we have developed (Walker, *et al.*, 2004).

The structure of the molecule is shown in Figure 1. The acetyl group is nearly coplanar with the phenyl group, the C2—C1—C7—O2 torsion angle being -1.64 (15)°. The NH(CH₂)₂COOH substituent is extended, with torsion angles C1—C2—N1—C9 - 176.52 (9), C2—N1—C9—C10 - 172.10 (9), N1—C9—C10—C11 179.81 (8), and C9—C10—C11—O2 173.73 (8)°. Thus, the molecule does not differ greatly from planarity, with mean and maximum deviations given in the abstract. The geometry of the intramolecular hydrogen bond, graph set S(6), (Etter, 1990) is given in Table 1, and is quite similar to that found in *N*-(2-acetylphenyl)acetamide (Slater *et al.*, 2006). The COOH group forms a typical hydrogen bonded dimer of graph set $R^2_2(8)$ about an inversion center, thus there are no extended networks of traditional hydrogen bonds.

Experimental

A round bottom flask containing acrylic acid (1.051 g, 14.58 mmol) was stoppered and placed into the refrigerator overnight. In a similar round bottom flask, 2'-aminoacetophenoneHCl (1.0056 g, 0.79199 g of liberated 2-aminoacetophenone 5.861 mmol) was dissolved in 10 ml of deionized water and chilled overnight. The round bottom flask containing the cold acrylic acid was then placed into to an ice bath to maintain a temperature of approximately 273 K. The ice bath was used because the reaction of acrylic acid and 2'-aminoacetophenone was thought to be exothermic. The cold 2'-aminoacetophenone solution was slowly added with stirring to the cold acrylic acid. The mixture was allowed to sit in the ice bath for 30 minutes and

supplementary materials

then gradually warmed to room temperature, then kept at room temperature with continuous stirring overnight. A yellow precipitate was isolated by gravimetric filtration and washed with deionized water. The yellow solid material resulted in a yield of 0.8963 g (79.991%).

Refinement

H atoms on C were placed in idealized positions with C—H distances 0.95 - 0.99 Å and thereafter treated as riding. Coordinates for the H atoms on N and O were refined. U_{iso} for H was assigned as 1.2 times U_{eq} of the attached atoms (1.5 for methyl and OH). A torsional parameter was refined for the methyl group.

Figures

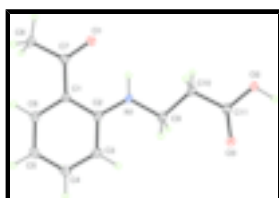


Fig. 1. A view of the title compound; ellipsoids have been plotted at the 50% level and H atoms have been assigned arbitrary radii.

3-(2-Acetylanilino)propanoic acid

Crystal data

$\text{C}_{11}\text{H}_{13}\text{NO}_3$

$M_r = 207.22$

Triclinic, $P\bar{1}$

Hall symbol: -P 1

$a = 5.1935(10)$ Å

$b = 9.8342(16)$ Å

$c = 9.920(2)$ Å

$\alpha = 77.084(12)^\circ$

$\beta = 85.174(11)^\circ$

$\gamma = 83.019(12)^\circ$

$V = 489.32(16)$ Å³

$Z = 2$

$F_{000} = 220$

$D_x = 1.406$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 2652 reflections

$\theta = 2.5\text{--}30.8^\circ$

$\mu = 0.10$ mm⁻¹

$T = 90.0(5)$ K

Fragment, colorless

$0.30 \times 0.20 \times 0.10$ mm

Data collection

Nonius KappaCCD diffractometer with an Oxford Cryosystems Cryostream cooler

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 90.0(5)$ K

ω scans with κ offsets

Absorption correction: none

11014 measured reflections

3012 independent reflections

2467 reflections with $I > 2s(I)$

$R_{\text{int}} = 0.021$

$\theta_{\text{max}} = 30.8^\circ$

$\theta_{\text{min}} = 2.6^\circ$

$h = -7 \rightarrow 7$

$k = -13 \rightarrow 13$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.041$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.113$	$w = 1/[\sigma^2(F_o^2) + (0.0573P)^2 + 0.1339P]$
$S = 1.04$	where $P = (F_o^2 + 2F_c^2)/3$
3012 reflections	$(\Delta/\sigma)_{\max} < 0.001$
143 parameters	$\Delta\rho_{\max} = 0.39 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\min} = -0.29 \text{ e } \text{\AA}^{-3}$
	Extinction correction: none

Special details

Experimental. ^1H (CDCl_3): δ (p.p.m.) 9.02 (s, 1H), 7.78 (d, $J = 8.0$ Hz, 1H), 7.38 (t, $J = 8.0$ Hz, 1H), 6.74 (d, $J = 8.0$ Hz, 1H), 6.62 (t, $J = 8.0$ Hz, 1H), 3.58 (t, $J = 7.0$ Hz, 2H), 2.73 (t, $J = 7.0$ Hz, 2H), and 2.60 (s, 3H). ^{13}C (CDCl_3): δ (p.p.m.) 201.1, 176.7, 150.5, 135.2, 132.9, 117.9, 114.5, 111.3, 37.9, 33.8, and 27.9. IR (thin film, KBr plates, cm^{-1}): 3297 (*m*), 2925 (*w*, *br*), 2881 (*sh*, *w*), 2635 (*w*), 2373 (*w*), 1698 (*s*), 1634 (*s*), 1569 (*m*), 1517 (*m*), 1430 (*m*), 1366 (*sh*), 1326(*m*), 1240 (*s*), 1156 (*w*), 1092(*w*), 1031(*w*), 943 (*m*), 825(*w*), 742 (*m*), 675(*w*), and 624 (*w*). Crystals were grown by evaporation from CDCl_3 .

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R - factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.96261 (15)	0.13504 (8)	0.35274 (7)	0.01656 (17)
O2	0.30686 (15)	-0.07587 (8)	0.93327 (7)	0.01593 (17)
H2O	0.183 (3)	-0.0849 (15)	1.0039 (16)	0.024*
O3	0.04143 (14)	0.11745 (8)	0.85415 (7)	0.01599 (17)
N1	0.53905 (17)	0.21868 (9)	0.49711 (8)	0.01437 (18)
H1N	0.673 (3)	0.1541 (14)	0.4924 (13)	0.017*
C1	0.70722 (18)	0.35189 (10)	0.27655 (10)	0.01194 (19)
C2	0.51897 (19)	0.33391 (10)	0.39115 (10)	0.01242 (19)
C3	0.3083 (2)	0.43948 (11)	0.39170 (10)	0.0161 (2)
H3	0.1787	0.4282	0.4657	0.019*
C4	0.2866 (2)	0.55873 (11)	0.28701 (11)	0.0168 (2)

supplementary materials

H4	0.1438	0.6283	0.2906	0.020*
C5	0.4725 (2)	0.57819 (11)	0.17590 (11)	0.0165 (2)
H5	0.4581	0.6605	0.1043	0.020*
C6	0.6778 (2)	0.47510 (10)	0.17238 (10)	0.0145 (2)
H6	0.8038	0.4879	0.0967	0.017*
C7	0.92650 (19)	0.24478 (10)	0.26361 (10)	0.01281 (19)
C8	1.1133 (2)	0.26868 (11)	0.13707 (10)	0.0157 (2)
H8A	1.2116	0.3472	0.1386	0.024*
H8B	1.0155	0.2905	0.0532	0.024*
H8C	1.2337	0.1838	0.1374	0.024*
C9	0.34277 (19)	0.19459 (11)	0.61054 (10)	0.0142 (2)
H9A	0.1798	0.1767	0.5757	0.017*
H9B	0.3060	0.2789	0.6508	0.017*
C10	0.43852 (19)	0.06928 (10)	0.72115 (10)	0.01378 (19)
H10A	0.4756	-0.0145	0.6801	0.017*
H10B	0.6025	0.0875	0.7548	0.017*
C11	0.24229 (19)	0.04048 (10)	0.84099 (10)	0.01244 (19)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0166 (4)	0.0146 (3)	0.0149 (3)	0.0019 (3)	0.0026 (3)	0.0012 (3)
O2	0.0148 (4)	0.0162 (4)	0.0127 (3)	0.0014 (3)	0.0038 (3)	0.0020 (3)
O3	0.0152 (4)	0.0162 (4)	0.0132 (3)	0.0019 (3)	0.0041 (3)	0.0003 (3)
N1	0.0136 (4)	0.0144 (4)	0.0111 (4)	0.0025 (3)	0.0048 (3)	0.0014 (3)
C1	0.0110 (4)	0.0123 (4)	0.0114 (4)	-0.0009 (3)	0.0016 (3)	-0.0012 (3)
C2	0.0127 (4)	0.0134 (4)	0.0104 (4)	-0.0015 (3)	0.0016 (3)	-0.0018 (3)
C3	0.0145 (5)	0.0179 (5)	0.0141 (4)	0.0014 (4)	0.0030 (3)	-0.0027 (4)
C4	0.0160 (5)	0.0150 (5)	0.0177 (5)	0.0030 (4)	0.0004 (4)	-0.0030 (4)
C5	0.0174 (5)	0.0135 (4)	0.0162 (5)	0.0002 (4)	0.0000 (4)	0.0005 (3)
C6	0.0152 (4)	0.0145 (4)	0.0125 (4)	-0.0021 (4)	0.0021 (3)	-0.0006 (3)
C7	0.0115 (4)	0.0146 (4)	0.0118 (4)	-0.0016 (3)	0.0015 (3)	-0.0024 (3)
C8	0.0146 (4)	0.0162 (5)	0.0136 (4)	0.0001 (4)	0.0045 (3)	-0.0005 (3)
C9	0.0126 (4)	0.0169 (5)	0.0109 (4)	-0.0001 (3)	0.0033 (3)	-0.0008 (3)
C10	0.0126 (4)	0.0155 (4)	0.0113 (4)	-0.0007 (3)	0.0028 (3)	-0.0005 (3)
C11	0.0132 (4)	0.0130 (4)	0.0107 (4)	-0.0022 (3)	0.0015 (3)	-0.0020 (3)

Geometric parameters (\AA , $^\circ$)

O1—C7	1.2384 (12)	C4—H4	0.9500
O2—C11	1.3228 (12)	C5—C6	1.3829 (14)
O2—H2O	0.909 (16)	C5—H5	0.9500
O3—C11	1.2278 (12)	C6—H6	0.9500
N1—C2	1.3628 (12)	C7—C8	1.5135 (13)
N1—C9	1.4519 (12)	C8—H8A	0.9800
N1—H1N	0.887 (14)	C8—H8B	0.9800
C1—C6	1.4075 (13)	C8—H8C	0.9800
C1—C2	1.4299 (13)	C9—C10	1.5193 (14)
C1—C7	1.4722 (14)	C9—H9A	0.9900

C2—C3	1.4145 (14)	C9—H9B	0.9900
C3—C4	1.3819 (14)	C10—C11	1.4993 (13)
C3—H3	0.9500	C10—H10A	0.9900
C4—C5	1.3983 (15)	C10—H10B	0.9900
C11—O2—H2O	108.7 (9)	O1—C7—C8	118.76 (9)
C2—N1—C9	122.44 (8)	C1—C7—C8	119.13 (8)
C2—N1—H1N	117.8 (9)	C7—C8—H8A	109.5
C9—N1—H1N	119.7 (9)	C7—C8—H8B	109.5
C6—C1—C2	118.53 (9)	H8A—C8—H8B	109.5
C6—C1—C7	119.47 (9)	C7—C8—H8C	109.5
C2—C1—C7	122.00 (8)	H8A—C8—H8C	109.5
N1—C2—C3	120.39 (9)	H8B—C8—H8C	109.5
N1—C2—C1	121.56 (9)	N1—C9—C10	109.68 (8)
C3—C2—C1	118.05 (9)	N1—C9—H9A	109.7
C4—C3—C2	121.48 (9)	C10—C9—H9A	109.7
C4—C3—H3	119.3	N1—C9—H9B	109.7
C2—C3—H3	119.3	C10—C9—H9B	109.7
C3—C4—C5	120.76 (9)	H9A—C9—H9B	108.2
C3—C4—H4	119.6	C11—C10—C9	111.66 (8)
C5—C4—H4	119.6	C11—C10—H10A	109.3
C6—C5—C4	118.62 (9)	C9—C10—H10A	109.3
C6—C5—H5	120.7	C11—C10—H10B	109.3
C4—C5—H5	120.7	C9—C10—H10B	109.3
C5—C6—C1	122.54 (9)	H10A—C10—H10B	107.9
C5—C6—H6	118.7	O3—C11—O2	122.63 (9)
C1—C6—H6	118.7	O3—C11—C10	123.78 (9)
O1—C7—C1	122.11 (9)	O2—C11—C10	113.59 (8)
C9—N1—C2—C3	3.20 (15)	C2—C1—C6—C5	-0.84 (15)
C9—N1—C2—C1	-176.52 (9)	C7—C1—C6—C5	178.13 (10)
C6—C1—C2—N1	-178.48 (9)	C6—C1—C7—O1	179.42 (9)
C7—C1—C2—N1	2.57 (15)	C2—C1—C7—O1	-1.64 (15)
C6—C1—C2—C3	1.80 (14)	C6—C1—C7—C8	-0.98 (14)
C7—C1—C2—C3	-177.15 (9)	C2—C1—C7—C8	177.96 (9)
N1—C2—C3—C4	178.57 (10)	C2—N1—C9—C10	-172.10 (9)
C1—C2—C3—C4	-1.71 (15)	N1—C9—C10—C11	179.81 (8)
C2—C3—C4—C5	0.58 (16)	C9—C10—C11—O3	-6.99 (14)
C3—C4—C5—C6	0.45 (16)	C9—C10—C11—O2	173.73 (8)
C4—C5—C6—C1	-0.30 (16)		

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O2—H2O \cdots O3 ⁱ	0.909 (16)	1.747 (16)	2.6531 (11)	174.7 (14)
N1—H1N \cdots O1	0.887 (14)	1.980 (14)	2.6690 (12)	133.4 (11)

Symmetry codes: (i) $-x, -y, -z+2$.

Fig. 1

